

# Engineering Chemistry 1 By Dr Ravikrishnan

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Engineering Chemistry 1 0CHE12 Page 3 When the circuit is completed the following reaction takes place At zinc electrode, Zn goes into the solution as Zn<sup>2+</sup> ions liberating electrons Zn Zn<sup>2+</sup> + 2e Oxidation half reaction At copper electrode Cu<sup>2+</sup> ions in electrolytic solution gets reduced to copper by gaining electrons from Zn

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a) Iron does not rust when zinc coating is broken in a galvanized iron pipe but rusting occurs if a coating

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